**Basics of Quantum Mechanics**

1. Which of these statements is true about the Bohr model and which is true of the Quantum mechanical model?
2. Electrons are treated as particles.
3. There are four variables that describe the energies of electrons
4. Depending on what type of orbital you have, different numbers of electrons can fit into a single orbital.
5. The farther an orbital is from the nucleus, the more energy the electrons in it have.
6. Electrons are 3-D waves centered on the nucleus.
7. It’s a mathematical model based on theory rather than pure observation.
8. The flame test can be explained both in terms of the Bohr model of the atom and the quantum model of the atom. How are these explanations similar, and how do they differ?
9. How does spectroscopy work?